

## Spin crossover in iron(II) complexes with 2,6-bis(1H-imidazole-2-yl)-4-chloropyridine

O. G. Shakirova,<sup>a,b\*</sup> T. D. Morozova,<sup>b</sup> I. A. Os'kina,<sup>c</sup> E. V. Korotaev,<sup>a</sup> S. V. Trubina,<sup>a</sup>  
V. V. Zvereva,<sup>a</sup> V. V. Kriventsov,<sup>d</sup> S. A. Petrov,<sup>e</sup> S. G. Kozlova,<sup>a</sup> A. Ya. Tikhonov,<sup>c</sup>  
L. G. Lavrenova<sup>a\*</sup>

<sup>a</sup> A.V. Nikolaev Institute of Inorganic Chemistry,  
Siberian Branch of the Russian Academy of Sciences, Russian Federation,  
630090 Novosibirsk, ave. Akad. Lavrentiev, 3.

E-mail: shakirova\_olga@mail.ru; ludm@niic.nsc.ru

<sup>b</sup> Komsomolsk-na-Amure State University, Russian Federation,  
681013 Komsomolsk-on-Amur, ave. Lenin, 27.

<sup>c</sup> N.N. Vorozhtsov Novosibirsk Institute of Organic Chemistry,  
Siberian Branch of the Russian Academy of Sciences, Russian Federation,  
630090 Novosibirsk, ave. Acad. Lavrentiev, 9.

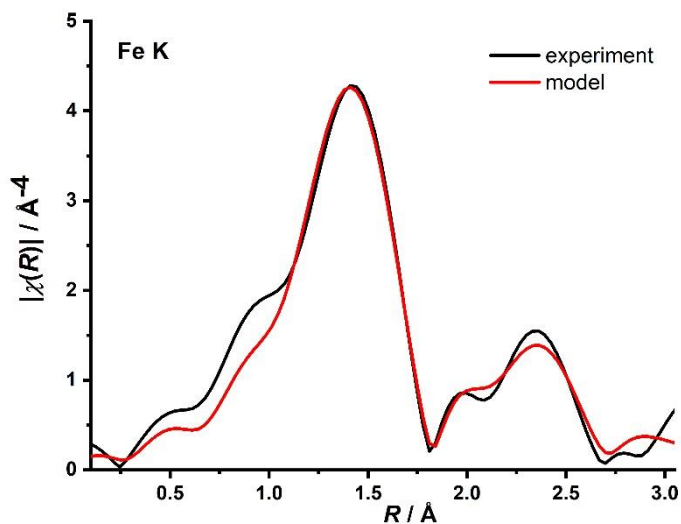
<sup>d</sup> Scientific Research Center SKIF G.K. Boreskov Institute of Catalysis, Siberian Branch of the  
Russian Academy of Sciences, Russian Federation,  
630559, Novosibirsk region, Naukograd Koltsovo, ave. Nikolsky, 1.

<sup>e</sup> Institute of Solid State Chemistry and Mechanochemistry,  
Siberian Branch of the Russian Academy of Sciences, Russian Federation,  
630090 Novosibirsk, Kutateladze str., 18.

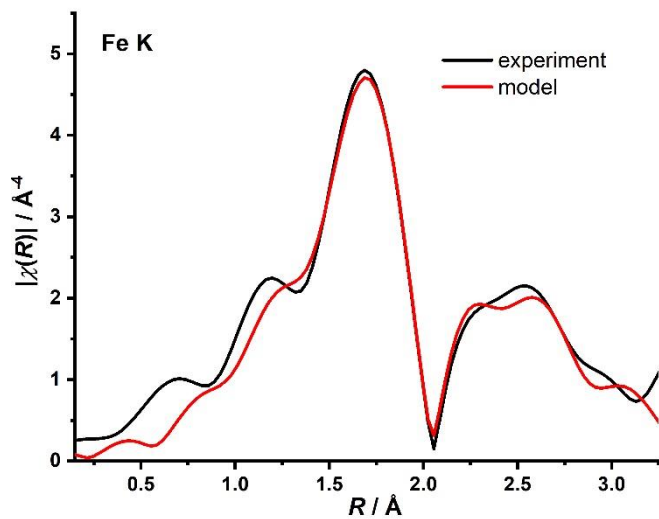
**Table 1S.** Parameters of the local atomic surrounding of Fe in coordination compounds. *R* is the interatomic distance, *N* is the coordination number,  $\sigma^2$  is the Debye-Waller factor, *R*-factor is the coefficient characterizing the quality of the fit

Compound		<i>R</i> / Å	<i>N</i>	$\sigma^2$ / Å <sup>2</sup>	$\Delta E_0$ / eV	<i>R</i> -factor
[FeL <sub>2</sub> ]SO <sub>4</sub> ·5H <sub>2</sub> O	Fe(1) – N(1)	1.95 ± 0.037	4.4 ± 0.97	0.008 ± 0.003	0.2 ± 5.3	0.0056
	Fe(1) – N(2)	2.14 ± 0.089	1.6 ± 0.35			
	Fe(1) – C(1)	2.85 ± 0.060	7 ± 1.8	0.015 ± 0.006		
	Fe(1) – C(2)	3.23 ± 0.074	5 ± 1.3			
[FeL <sub>2</sub> ](NO <sub>3</sub> ) <sub>2</sub> ·2H <sub>2</sub> O	Fe(1) – N(1)	1.99 ± 0.060	1.2 ± 0.61	0.006 ± 0.003	4.2 ± 3.1	0.0190
	Fe(1) – N(2)	2.16 ± 0.025	4.8 ± 2.4			
	Fe(1) – C(1)	3.02 ± 0.041	6 ± 1.6	0.009 ± 0.003		
	Fe(1) – C(2)	3.41 ± 0.032	6 ± 1.6			

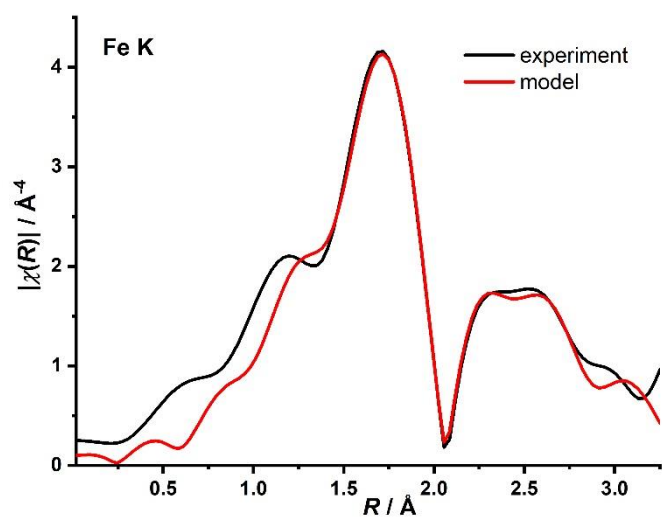
[FeL <sub>2</sub> ]Cl <sub>2</sub> ·3H <sub>2</sub> O	Fe(1) – N(1)	1.99 ± 0.048	1.3 ± 0.65	0.006 ± 0.003	4 ± 2.9	0.0200
	Fe(1) – N(2)	2.17 ± 0.024	4.7 ± 2.4			
	Fe(1) – C(1)	3.03 ± 0.040	6 ± 1.4	0.010 ± 0.003		
	Fe(1) – C(2)	3.41 ± 0.034	6 ± 1.4			



**Fig. 1S.** Fourier transformed Fe K-edge EXAFS spectrum (without phase correction) of [FeL<sub>2</sub>]SO<sub>4</sub>·5H<sub>2</sub>O ( $R$ -range = 1.0 – 3.0 Å,  $k$ -range = 3.0-11.0 Å<sup>-1</sup>).



**Fig. 2S.** Fourier transformed Fe K-edge EXAFS spectrum (without phase correction) of [FeL<sub>2</sub>](NO<sub>3</sub>)<sub>2</sub>·2H<sub>2</sub>O ( $R$ - range = 1.0 – 3.2 Å,  $k$ - range = 2.8-11.2 Å<sup>-1</sup>).



**Fig. 3S.** Fourier transformed Fe K-edge EXAFS spectrum (without phase correction) of  $[\text{FeL}_2]\text{Cl}_2 \cdot 3\text{H}_2\text{O}$  ( $R$ - range = 1.0 – 3.2 Å,  $k$ - range = 2.8-11.2 Å<sup>-1</sup>).